

SNS COLLEGE OF TECHNOLOGY

(An Autonomous Institution) COIMBATORE-35.



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DEPARTMENT OF AUTOMOBILE ENGINEERING

COURSE NAME : 19AUT202 - HYBRID ELECTRIC & FUEL CELL VEHICLE

II YEAR /III SEMESTER

Unit 4- Introduction to Fuel Cells

Topic : Molten Carbonate Fuel Cell



INTRODUCTION



- Molten-carbonate fuel cells are high-temperature fuel cells that operate at temperatures of 600 °C and above.
- Molten carbonate fuel cells are currently being developed for natural gas and coalbased power plants for electrical utility, industrial, and military applications.





SPECIFICATIONS



- Fuel Hydrogen
- Oxidant Oxygen
- Catalyst Nickel
- > Electrolyte Alkali (Li, Na, and K) carbonates stabilized in an LiAlO2 ceramic matrix





CONSTRUCTION



- MCFC's use a liquid electrolyte (molten carbonate) which consists of a sodium (Na) and potassium (K) carbonate.
- This electrolyte is supported by a ceramic (LiAlO₂) matrix to contain the liquid between the electrodes.
- > Common MCFC electrolytes contain 62% Li_2CO_3 and 38% K_2CO_3 .



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CONSTRUCTION



- > The anode material typically consists of a porous Ni based alloy.
- ➤ Ni is alloyed with either Chromium or Aluminum in the 2-10% range.
- On the other side of the cell, the cathode material is composed of either Lithium metatitanate or of a porous Ni that is converted to a lithiated nickel oxide.





FUEL CELL WORKING







CHEMICAL REACTION



> Anode Reaction:

 $\mathrm{H_2} + \mathrm{CO_3}^{2\text{-}} \rightarrow \mathrm{H_2O} + \mathrm{CO_2} + 2\mathrm{e}^{-1}$

Cathode Reaction:

 $\frac{1}{2}O_2 + CO_2 + 2e^- \rightarrow CO_3^{2-}$

> Overall Reaction:

 $\mathrm{H}_2 + \frac{1}{2} \mathrm{O}_2 \to \mathrm{H}_2\mathrm{O}$



APPLICATIONS



Used in Large stationary Power generation





REFERENCE



https://www.sciencedirect.com/topics/engineering/molten-carbonate-fuel-cell
https://en.wikipedia.org/wiki/Molten_carbonate_fuel_cell







THANK YOU !!!

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