



## Unit II - Topic 3

### Arrhenius equation - Thermal Death Time Curves – Optimum temperature

The Arrhenius equation is an expression that provides a relationship between the rate constant (of a chemical reaction), the absolute temperature, and the A factor (also known as the pre-exponential factor; can be visualized as the frequency of correctly oriented collisions between reactant particles). It provides insight into the dependence of reaction rates on the absolute temperature.

What is the Arrhenius Equation?

The expression of the Arrhenius equation is:



$$k = Ae^{-E_a/RT}$$

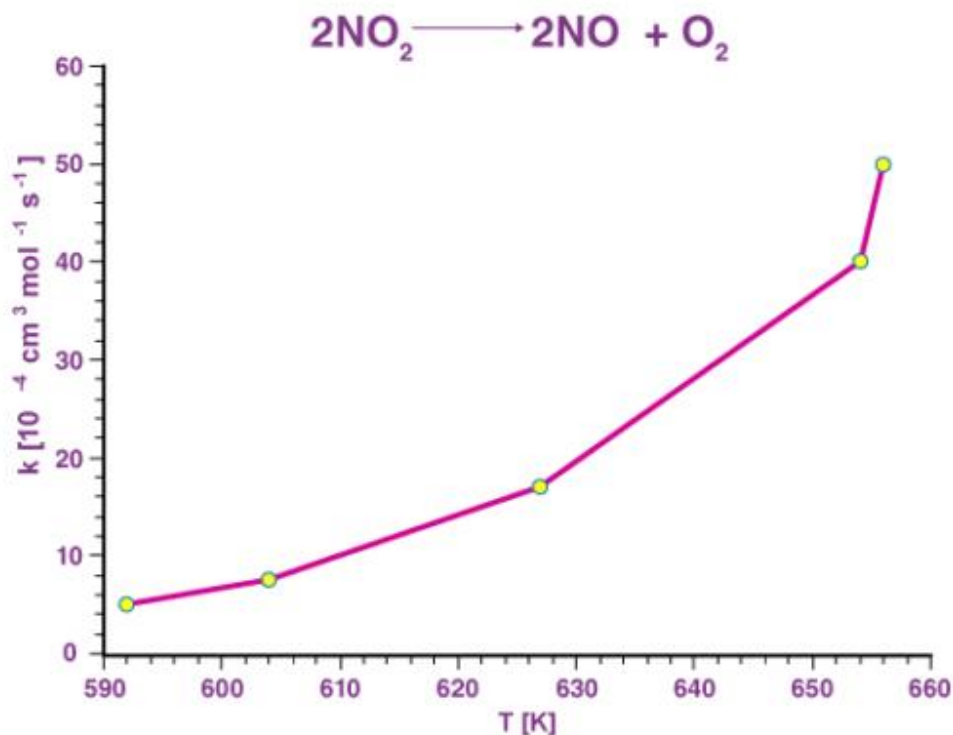
Where,

- k denotes the rate constant of the reaction
- A denotes the pre-exponential factor which, in terms of the collision theory, is the frequency of correctly oriented collisions between the reacting species
- e is the base of the natural logarithm (Euler's number)
- $E_a$  denotes the activation energy of the chemical reaction (in terms of energy per mole)
- R denotes the universal gas constant
- T denotes the absolute temperature associated with the reaction (in Kelvin)

If the activation energy is expressed in terms of energy per reactant molecule, the universal gas constant must be replaced with the Boltzmann constant ( $k_B$ ) in the Arrhenius equation. The Arrhenius equation was put forward by the Swedish chemist Svante Arrhenius in the year 1889.

#### Graphical Representation of the Arrhenius Equation

For the decomposition reaction undergone by nitrogen dioxide (given by  $2\text{NO}_2 \rightarrow 2\text{NO} + \text{O}_2$ ), a graph plotted with the rate constant (k) on the Y-axis and the absolute temperature (T) on the X-axis is provided below. Note that the rate of the reaction increases as the temperature increases.



### Arrhenius Plot

When logarithms are taken on both sides of the equation, the Arrhenius equation can be written as follows:

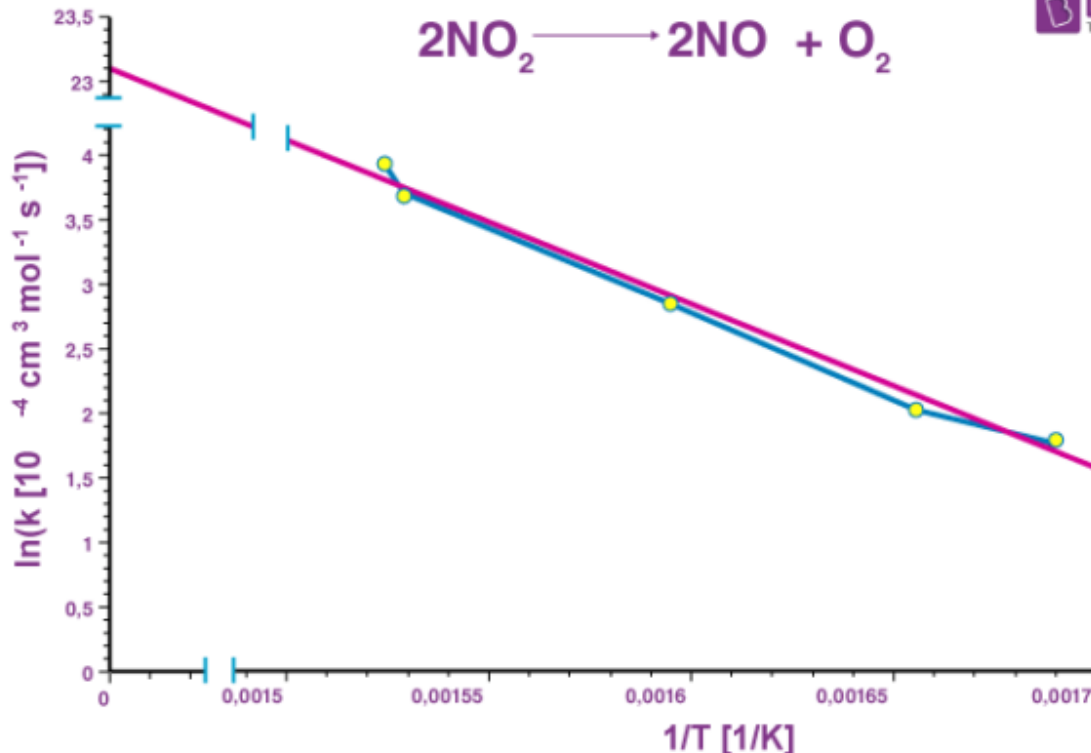
$$\ln k = \ln(Ae^{-E_a/RT})$$

Solving the equation further:

$$\ln k = \ln(A) + \ln(e^{-E_a/RT})$$

$$\ln k = \ln(A) + (-E_a/RT) = \ln(A) - (E_a/R)(1/T)$$

Since  $\ln(A)$  is a constant, the equation corresponds to that of a straight line ( $y = mx + c$ ) whose slope ( $m$ ) is  $-E_a/R$ . When the logarithm of the rate constant ( $\ln K$ ) is plotted on the Y-axis and the inverse of the absolute temperature ( $1/T$ ) is plotted on the X-axis, the resulting graph is called an Arrhenius plot.



The Arrhenius plot for the decomposition of nitrogen dioxide is illustrated above.

Does the Arrhenius Equation Account for Catalysts?

The function of a catalyst is to lower the activation energy required by a reaction. Therefore, the lowered activation energy (accounted by the catalyst) can be substituted into the Arrhenius equation in order to obtain the rate constant for the catalysed reaction.

The exponential part of the Arrhenius equation ( $-E_a/RT$ ) accounts for an exponential increase in the value of the rate constant for any decrease in the activation energy. Since the rate of a chemical reaction is directly proportional to the rate constant of that reaction, the decrease in activation energy results in an exponential increase in the reaction rate.

It is important to note that the rates of uncatalysed reactions are more severely affected by temperature than the rates of catalysed reactions. This is because the activation energy is in the numerator of the exponential term  $-E_a/RT$  and the absolute temperature is in the denominator. Since the activation energy of the catalysed reaction is relatively low, the effect of temperature on the rate constant is more visible in the corresponding uncatalysed reaction.

Solved Exercises

Example 1

**The activation energy of a chemical reaction is 100 kJ/mol and its A factor is  $10 \text{ M}^{-1}\text{s}^{-1}$ . Find the rate constant of this equation at a temperature of 300 K.**



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Given,

$$E_a = 100 \text{ kJ.mol}^{-1} = 100000 \text{ J.mol}^{-1}$$

$$A = 10 \text{ M}^{-1}\text{s}^{-1}, \ln(A) = 2.3 \text{ (approx.)}$$

$$T = 300 \text{ K}$$

The value of the rate constant can be obtained from the logarithmic form of the Arrhenius equation, which is:

$$\ln k = \ln(A) - (E_a/RT)$$

$$\ln k = 2.3 - (100000 \text{ J.mol}^{-1})/(8.314 \text{ J.mol}^{-1}.\text{K}^{-1})*(300\text{K})$$

$$\ln k = 2.3 - 40.1$$

$$\ln k = -37.8$$

$k = 3.8341*10^{-17} \text{ M}^{-1}\text{s}^{-1}$  (from the units of the A factor, it can be understood that the reaction is a second-order reaction, for which the unit of k is  $\text{M}^{-1}\text{s}^{-1}$ )

Therefore, the value of the rate constant for the reaction at a temperature of 300K is approximately  $3.8341*10^{-17} \text{ M}^{-1}\text{s}^{-1}$ .

Example 2

**At a temperature of 600 K, the rate constant of a chemical reaction is  $2.75*10^{-8} \text{ M}^{-1}\text{s}^{-1}$ . When the temperature is increased to 800K, the rate constant for the same reaction is  $1.95*10^{-7}\text{M}^{-1}\text{s}^{-1}$ . What is the activation energy of this reaction?**

Given,

$$T_1 = 600\text{K}$$

$$k_1 = 2.75*10^{-8} \text{ M}^{-1}\text{s}^{-1}.$$

$$T_2 = 800\text{K}$$

$$K_2 = 1.95*10^{-7}\text{M}^{-1}\text{s}^{-1}$$

When the A factor is eliminated from the Arrhenius equation, the following equation is obtained:  
 $\ln(k_1/k_2) = (-E_a/R)(1/T_1 - 1/T_2)$

Substituting the given values in the equation, the value of  $E_a$  can be determined:

$$\ln(2.75*10^{-8}/1.95*10^{-7}) = (-E_a/8.314 \text{ J.K}^{-1}.\text{mol}^{-1})*(0.00041\text{K}^{-1})$$

$$\ln(0.141) = (E_a)*(-0.0000493) \text{ J}^{-1}.\text{mol}$$

$$E_a = (-1.958)/(-0.0000493)\text{J.mol}^{-1} = 39716 \text{ J.mol}^{-1}$$

The activation energy of the reaction is approximately  $39716 \text{ J.mol}^{-1}$ .



## Frequently Asked Questions – FAQs

What is Arrhenius equation explain with example?

The Arrhenius equation is sometimes expressed as  $k = Ae^{-E/RT}$  where  $k$  is the rate of chemical reaction,  $A$  is a constant depending on the chemicals involved,  $E$  is the activation energy,  $R$  is the universal gas constant, and  $T$  is the temperature.

Which theory is the basis for Arrhenius equation?

The number of effective collisions is proportional to the number of molecules above certain threshold energy. As the temperature increases, so does the number of molecules with energies exceeding the threshold energy.

What is the effect of catalyst on the rate of reaction?

A catalyst decreases the activation energy barrier of the reaction, thereby increasing the rate of the reaction.

What is the value of  $A$  and  $E_a$  in Arrhenius equation?

In the Arrhenius equation for a certain reaction, the value of  $A$  and  $E_a$  (energy of activation) are  $4 \times 10^{-13} \text{sec}^{-1}$  and  $98.6 \text{ kJ mol}^{-1}$  respectively.

What is true about the Arrhenius equation?

The exponential term in the Arrhenius equation implies that the rate constant of a reaction increases exponentially when the activation energy decreases. Because the rate of a reaction is directly proportional to the rate constant of a reaction, the rate increases exponentially as well.

To learn more about the Arrhenius equation and other important concepts related to chemical kinetics, register with BYJU'S and download the mobile application on your smartphone.

### **Thermal Death Time (TDT)**

It is defined as number of minutes of heating required to destroy a specified number of organisms at a given temperature. The thermal death time follows a logarithmic relationship with the temperature of heating. This means as we increase or decrease the temperature of heating, the thermal death time for a bacteria will increase or decrease in a logarithmic scale. Therefore if we plot the TDT values on a log scale of a semilog paper (The graph paper which has vertical grids on log scale and horizontal grids on linear scale), and the temperature of heating on linear scale, we would get a straight line. When TDT plotted on the log scale and temperature of heating on linear scale, a straight line of best fit is known as thermal destruction curve or thermal death time (TDT) curve. The TDT curve constructed in this manner is called end point curve. This curve can be used for finding out the thermal death time for a fixed number of organisms at an unknown temperature. TDT values obtained in this manner are valid only for the number and type of organisms used in that particular experiment. TDT curves for the same organism present at higher density would be different. Instead of using semilog paper, TDT curve can also be



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plotted on a linear graph paper by taking log of the TDT values on “Y” axis and heating temperature on X axis. While D values represent killing of 90% of a population and TDT represent destruction of 100% of the population they are always proportional.