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Pharmaceutical Inorganic chemistry

Question bank

UNIT-2

Chapter: Acid-Base and Buffers

2marks

1. Define the following terms i) Osmotic pressure ii) Isotonic solution.
2. What is buffer capacity?
3. Define buffers with examples.
4. Define buffers. Give examples of two official buffers.
5. What are buffer capacity and isotonicity?
6. What is the importance of buffer in pharmacy?
7. Define the term hypotonic and hypertonic.
8. Define the term tonicity.
9. What is osmotic pressure?
10. Give any two examples of strong acid and weak acid.
11. Give any two examples of a strong base and weak base.
12. Write the limitations of Arrhenius theory.

5marks

1. What are buffers derive from the Henderson-hasselbalch equation for buffers?
2. What are buffered isotonic solutions? Give details.
3. Explain the Lewis acid and Lewis base with examples.
4. What is a buffer solution that explains the mechanism of buffer action?
5. Write the importance of buffer solutions in pharmacy.

6. Write a note on the isotonic buffer.
7. Define acids and bases according to various concepts.
8. Write a note on buffer solutions.
9. Describe various methods used to adjust isotonicity.

10Marks

1. Describe buffer capacity, stability of buffers, methods of adjusting isotonicity. Buffers and their role in pharmacy.
2. Discuss various types of physiological buffers. Explain the mechanism of their Buffer action.
3. Explain major buffers used in pharmaceutical preparations. What are the disadvantages of each? What factors must be taken into consideration in the Selection of a buffer?
4. What are buffers? Explain the mechanism of buffer action with an example. Briefly discuss the role of buffers in pharmacy.