

BP301T.

PHARMACEUTICAL ORGANIC CHEMISTRY –II (Theory)

UNIT – V CYCLO-ALKANES

What are cyclo-alkanes?

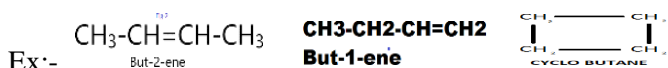
Cyclo-alkanes are also called cycloparaffins.

They are Alicyclic Compounds i.e. they are both aliphatic and cyclic compounds, and their properties are almost similar to alkanes.

Cyclo-alkanes are saturated hydrocarbons, that means all the carbons present do not contain any pi bond, all their valency are satisfied with sigma bonds.

Cyclo-alkanes have General Formula C_nH_{2n} with 1st member Cyclo-propane (Formula: C_3H_6).

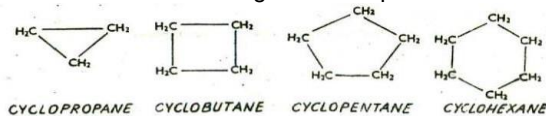
Isomerism: - Cycloalkanes have same molecular formula to that of corresponding Alkenes. So they are functional isomers with alkenes of same carbon number.



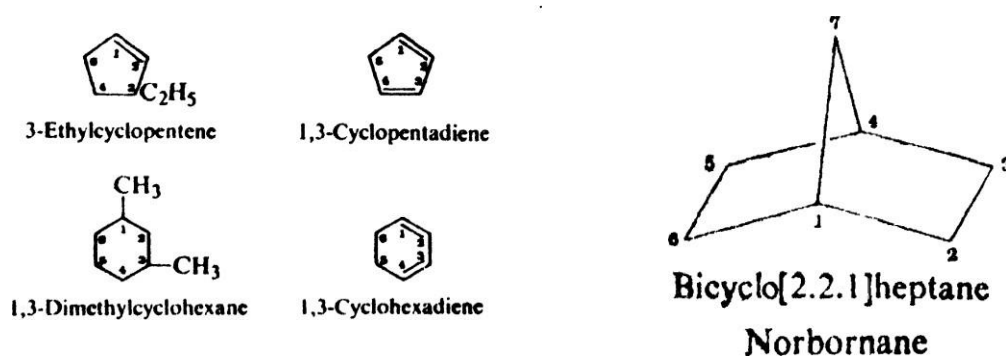
All the above three structures have molecular formula C_4H_8 , But 1st two figures are functional isomer with the third figure.

Nomenclature: It can be read in 3 sections

1. **Naming of simple cycloalkanes:** - It is done by attaching the prefix *cyclo* to the name of the corresponding normal alkane having the same number of carbon atoms as in the ring. For example

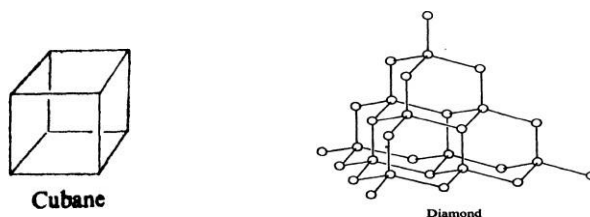


2. **Naming of substituted cycloalkanes:** - Naming of substituents in cycloalkanes as shown in 1st figure below are similar to that of straight chain compounds. But in 2nd compound which is a fused cycloalkane its typical.



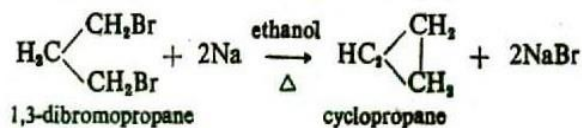
Lets brake the name Bicyclo[2.2.1]heptane. Bicyclo implies the compound has two rings i.e. if we break two carbon bonds, the compound will be an alkane of 7 carbon and can be a heptane. Now the most important part naming of '2.2.1'. It is so named because the number of carbons except bridgeheads and shared carbons is two (C-2 and C-3), two (C-5 and C-6), and the bridge head made of one carbon i.e. C-7.

3. **Example of typical cycloalkanes.:** Some typical fused cyclo-alkanes are given below. In cubane each carbon attached to three other carbon, in diamond each carbon attached to 4 other carbons.

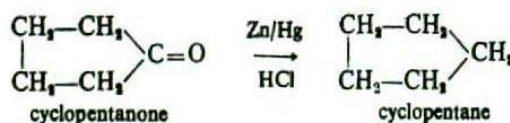
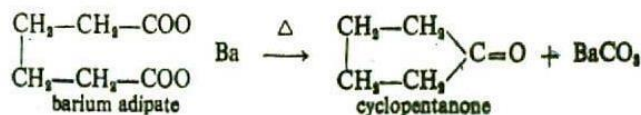


Methods of Preparation:

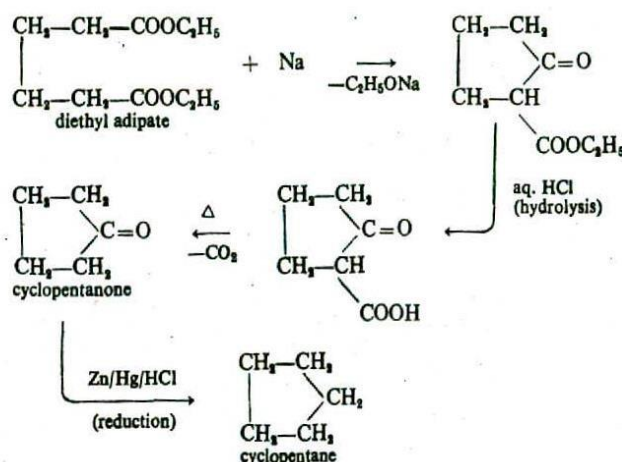
- From Di-halogen Compounds:** - Suitable 1,3 or 1,4 like di-halogen alkanes on treatment with sodium or zinc give corresponding cycloalkanes. For example



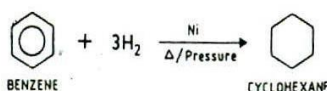
- From Calcium or Barium salts of Dicarboxylic acids:** - When the calcium or barium salt of adipic, pimelic, or suberic acid is heated, a cyclic ketone is formed. Cyclic ketones may be readily converted into the corresponding cycloalkanes by means of the *Clemmensen Reduction*.



- From Esters of Dicarboxylic acids (Dieckmann Reaction):** - The diester of adipic, pimelic, or suberic acid when treated with sodium undergoes intramolecular acetoacetic ester condensation and a keto-ester is formed. The keto-ester on hydrolysis give corresponding cyclic ketones. Cyclic ketones on reduction yield the corresponding cyclo- ketones.



- From Aromatic Compounds.** Benzene may be catalytically hydrogenated at elevated temperature and pressure to yield cyclohexane.

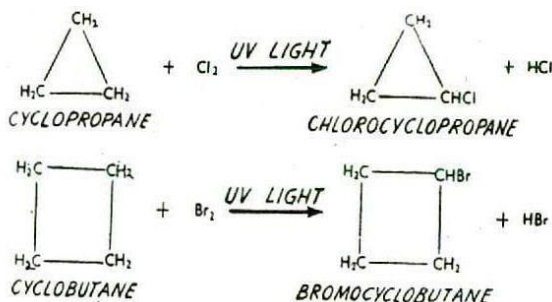


CHEMICAL REACTIONS

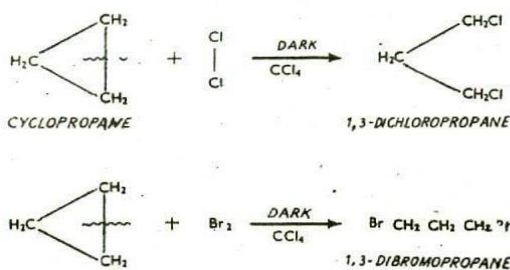
Cyclo-alkanes react mostly like their corresponding to their straight chain alkanes.

- Halogenation in presence of UV-light:-** Produce mono substituted halogen without breaking the ring.

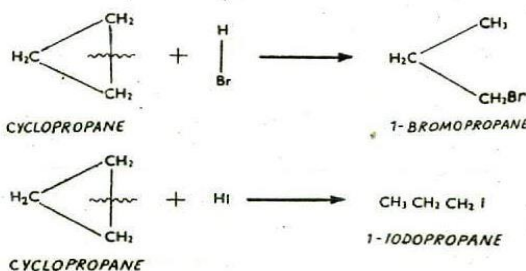
EX-



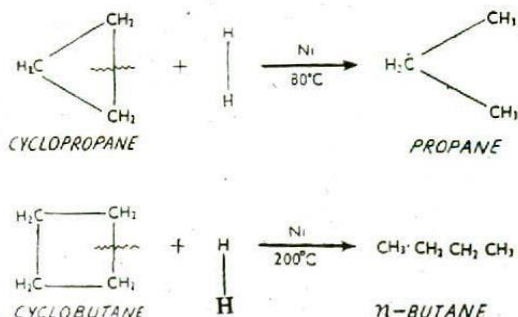
2. **Addition of Cl₂/ Br₂ in presence of CCl₄ in dark:-** Cyclopropane reacts with Cl₂/ Br₂ at room temperature and in the absence of diffused sunlight to produce 1,3-dichloropropane and 1,3-dibromopropane respectively. One of the carbon-carbon bonds of cyclopropane is broken and the two halogen atoms appear at the ends of the propane chain.



3. **Reaction with concentrated HBr and HI:-** Produce 1-bromopropane and 1-iodopropane respectively with ring breaking.



4. **Addition of Hydrogen / Catalytic Reduction:-** Cyclopropane and cyclobutane react with hydrogen in the presence of a nickel catalyst to produce n-propane and n-butane respectively. Cyclopentane and higher members of the family do not give this reaction.



STABILITY OF CYCLO-ALKANES

A stable organic compound means it is very less reactive. By the term stability we get an idea about chemical reactivity of a compound. To describe stability of cycloalkanes some very famous theoretical concepts are there viz.

Bayer Strain Theory

Sachse Mohr Concept

Coulson and Moffitt's modification

Bayer Strain Theory

Adolf von Baeyer, a German chemist and a Nobel prize winner from university of Munich, Germany in 1885 proposed a theory to explain the relative stability of the first few cycloalkane. The theory based on following facts.

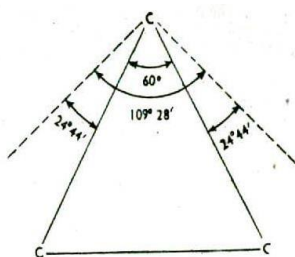
- . Cyclo-alkanes are saturated compounds. So all the carbons should have normal tetrahedral angle of 109.5° .
- . Any deviation of bond angles from the normal tetrahedral value would, impose a condition of internal strain on the ring called angle strain.
- . More the value of angle strain less stable the compound is.
- . He assumes that all the cyclo-alkanes are flat and coplanar that means they are two dimensional and present in one plane.



A.V. Baeyer

Case-Study:-

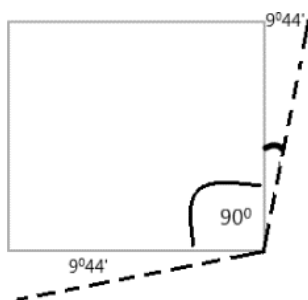
1. Cyclopropane:-



In cyclopropane, the three carbon atoms occupy the corners of an equilateral triangle. Thus cyclopropane has C—C—C bond angles of 60° . This implies that the normal tetrahedral angle of 109.5° between any two bonds is compressed to 60° , and that each of the two bonds involved is pulled in by $\frac{1}{2}(109^\circ 28' - 60^\circ) = 24^\circ 44'$.

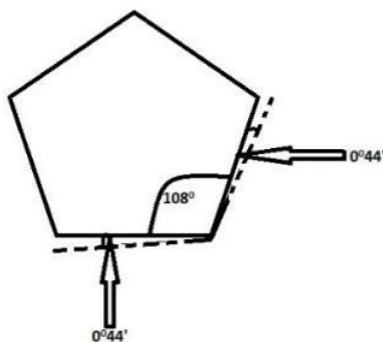
The value $24^\circ 44'$ then represents the *angle strain* or the deviation through which each bond bends from the normal tetrahedral direction.

2. Cyclo-butane:-



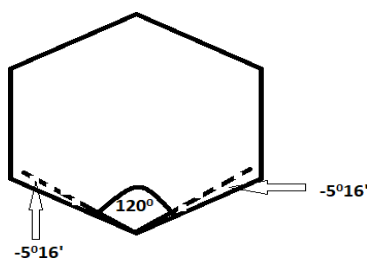
Here the four carbon atoms occupy the corners of a square. So it has C—C—C bond angles of 90° & angle strain on each bond is $\frac{1}{2}(109^\circ 28' - 90^\circ) = 9^\circ 44'$.

3. Cyclo-Pentane:-



Here the five carbon atoms lie at the corners of a regular pentagon. Thus cyclopentane has C—C—C bond angles of 108° . The angle strain which each bond fills is $1/2(109^\circ 28' - 108^\circ) = 0^\circ 44'$.

4. Cyclo-hexane:-



In cyclohexane the six carbon atoms occupy the corners of a regular hexagon. Thus cyclohexane has C—C—C bond

angles of 120° . The angle strain will be $1/2(109^\circ 28' - 120^\circ) = -5^\circ 16'$

Similarly, in the case of cyclo-heptane the deviation from the normal tetrahedral angle is $-9^\circ 33'$. In the case of cyclooctane it is $-12^\circ 46'$.

Derivation:-

Sl.No	Name	Angle strain
1	Cyclo-propane	$+24^\circ 44'$
2	Cyclo-butane	$+9^\circ 44'$
3	Cyclo-pentane	$+0^\circ 44'$
4	Cyclo-hexane	$-5^\circ 16'$
5	Cyclo-heptane	$-9^\circ 33'$
6	Cyclo-octane	$-12^\circ 46'$

1. The + sign indicates that the C—C bonds have to be compressed to satisfy the geometry of the ring. The - sign indicates that the C—C bonds have to be widened to satisfy the geometry of the ring. *Whether the angle strain is positive or negative, its magnitude determines the extent of strain in the ring.*
2. The deviation from the normal tetrahedral angle is *maximum* in the case of cyclopropane. So, it is most unstable and open up on the slightest provocation and thus releasing the strain within.
3. The deviation from the normal tetrahedral angle is *minimum* in the case of cyclopentane. This implies that cyclopentane is under least strain and should be most stable
4. According to the Baeyer Strain Theory, cyclohexane and the higher cycloalkanes should become increasingly unstable and hence more reactive.

Drawbacks of Bayer-strain theory

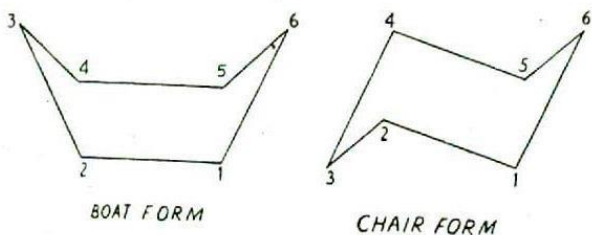
1. The limitation of Bayer strain theory is that he assumed that all cycloalkanes are planar.
2. Baeyer was unable to explain the effect of angle strain in Larger Ring System.
3. According to Baeyer Cyclopentane should be much stable than cyclohexane but practically it is reversed.
4. Larger ring systems are not possible according to Baeyer as they have negative strain but they exist and much stable. - Larger ring systems are not planar but puckered to eliminate angle strain.

SACHSE MOHR'S CONCEPT OF STRAINLESS RINGS

In order to account for the stability of cycloalkanes beyond cyclopentane, Sachse and *Mohr* (1918) pointed out that such rings can become absolutely free of strain if all the ring carbons are not forced into one plane as was supposed by Baeyer.

If the ring assumes a such 3D condition the normal tetrahedral angles of $109^{\circ}28'$ are retained and as a result, the strain within the ring is relieved. Thus, cyclohexane can exist in two non-planar strainless forms, namely, the Boat form and the Chair form.

In the Boat form, carbons 1, 2, 4 and 5 lie in the same plane and carbons 3 and 6 above the plane. In the chair form, carbons 1, 2, 4 and 5 lie in the same plane, but carbon 6 is above the plane and carbon 3 is below it.



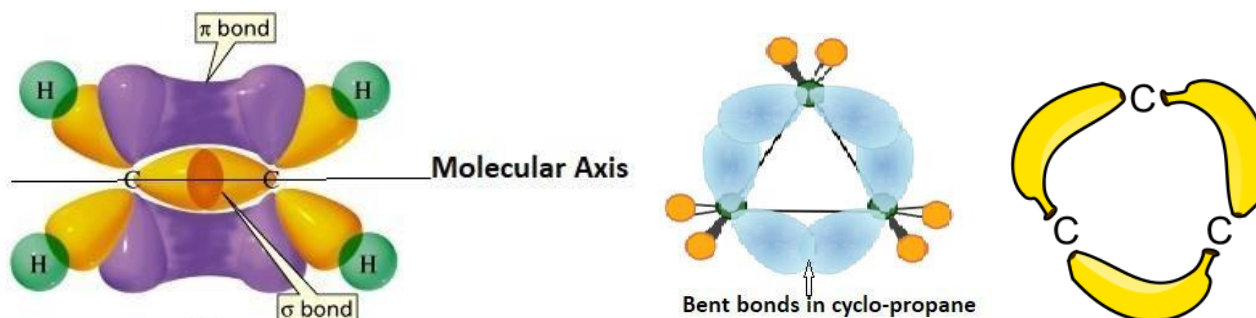
Actually, only one form of cyclohexane is known and not two forms as shown above. The failure to isolate the two forms is due to rapid interconversions between them. Such non-planar strainless rings in which the ring carbon atoms can have normal tetrahedral angles are also possible for larger ring compounds.

Coulson and Moffitt's modification (Bent bond/Banana bond)

Coulson-Moffitt modify Baeyer strain theory to show the actual position of carbon-carbon bonds in cyclopropane, which somewhat relieves the strain on C-C bonds.

So, to satisfy regular tetrahedral geometry ($109^{\circ}28'$) and equilateral triangle angle (60°) they introduce a new type of bond called banana bond/bent bond which is intermediate between sigma bond and pi-bond in case of overlapping.

What is this Bent bond:- Sigma bonds have head-on or end-to-end overlapping. The hybrid orbitals including the intersection bonding region are present equally on both side of the molecular axis. But in case of pi-bonds the lateral or sidewise overlapping of p-orbitals occur and here the bonding intersection area is totally out of the molecular axis.



In Bent-bond (usually they are sigma bonds) the hybrid orbitals and intersection bonding area are not present equally on the both side of molecular axis. The bonds are bent towards the direction of angle strain. So they are somewhat more unstable than sigma bonds and less unstable than pi bonds.

Important Questions often asked from the unit

1. Baeyer Strain theory. 10Marks
2. Sachse-Mohr's concept. 5 Marks
3. Coulson and Moffitt's modification. 5 Marks
4. What is bent bond. 2 marks
5. Method of preparation and chemical reactions for short questions.



REFERENCES

1. Organic Chemistry by Morrison and Boyd
 2. Arun Bhal. A Text Book of Organic Chemistry, 22nd Edition Page no: 618
 3. P L Soni. Text Book of Organic Chemistry, 29th Edition Page no: 2.58 & 2.477.
 4. https://profiles.uonbi.ac.ke/andakala/files/sch_102_lecture_week_8_2018-9.pdf
 5. http://yengage.yenepoya.edu.in/idata/YenepoyaUniversity/ilFile/4/59/file_45980/001/Unit%20V%20POC%20I%20Cycloalkane.pdf
-