



SNS COLLEGE OF PHARMACY AND HEALTH SCIENCES

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Expectorant s and Emetics

Expectorants

“Remove sputum from
respiratory tract.”

E.g.: Potassium chloride



Emetics

“Induce vomiting”

E.g.: Antimony potassium
tartrate



Ammonium chloride



Syn:

Ammonium muriate

MOP:

From ammonia and hydrochloric acid.



Physical Properties:

1. White crystalline powder
2. Odourless
3. Ammonical taste
4. Very soluble in water
5. Soluble in glycerin
6. Soluble in alcohol
7. Insoluble in acetone and ether
8. Hygroscopic in nature

Uses

1. Mild expectorants
2. To maintain acid base balance in body.
3. In cough preparation.



Storage

“ It should be stored in well closed container at a cool place.”



Incompatibility

Incompatible with alkalis,
carbonates, lead and silver salts.



Potassium Iodide

MOP:

By using iodine, ferric salts and potassium carbonate.



Physical Properties:

1. White crystalline powder
2. Odourless
3. Bitter taste
4. Soluble in water
5. Soluble in glycerin
6. Soluble in alcohol.



Uses

1. Mild expectorants
2. To treatment of hypo- thyroidism.
3. Used as a source of iodine.
4. In cough preparation.
5. To stabilize in preparation of solutions of iodine.



Storage

“ It should be stored in well closed container at a cool place.”



Incompatibility

Incompatible with salts of mercury, bismuth, lead, iron, oxidizing agent and dilute hydrochloric acid.



Antimony Potassium tartrate

[K O O C - C H O H - C H O H - C O O - S b O]

Or

C₄H₄KO₇Sb.

Syn:

Tartar of emetics, tartar emetic

MOP

By mixing antimony trioxide (5 parts) with potassium acid tartrate (6 parts).



Physical Properties:

- Colorless powder.
- Odourless.
- Sweet taste.
- Soluble in water.
- Insoluble in alcohol.
- Aqueous solution is acidic in nature.

Uses

1. As a emetics.
2. As a expectorants.
3. As a anti microbial agent.



Storage

“ It should be stored in well closed container at a cool place.”



Incompatibility

Incompatible with acid salts, heavy metals, albumin and soap solution.

