



19CH201 - ENGINEERING CHEMISTRY

UNIT-1 - ELECTROCHEMISTRY

1.7. Electrochemical cell -Galvanic cell

TYPES OF CELLS

A cell is a device consisting two half cell. Each half cell consists of an electrode dipped in an electrolytic solution. The two half cells are connected through one wire. The followings are two types of cells.

1. Electrolytic cells.
2. Electrochemical cells (or) voltaic cells (or) galvanic cells.

ELECTROLYTIC CELLS

Electrolytic cells are cells in which electrical energy is used to bring about the chemical reaction.

Ex: Electrolysis, electroplating, etc.,

ELECTROCHEMICAL CELLS (or) GALVANIC CELLS

Electrochemical cells are entirely different from electrolytic cells. The cells used for electrolysis (where electrical energy is converted to chemical energy) are called electrolytic cells, whereas in electrochemical cells, chemical energy is converted to electrical energy.

Galvanic cells are electrochemical cells in which the electrons, transferred due to redox reaction, are converted to electrical energy.

Ex: Daniel cell (Fig 1.7)

Cell device (Construction)

Daniel cell consists of a zinc electrode dipped in 1 M $ZnSO_4$ solution and a copper electrode dipped in 1 M $CuSO_4$ solution. Each electrode is known as a half cell. The two solutions are inter connected by a salt bridge and the two electrodes are connected by a wire through the voltmeter.

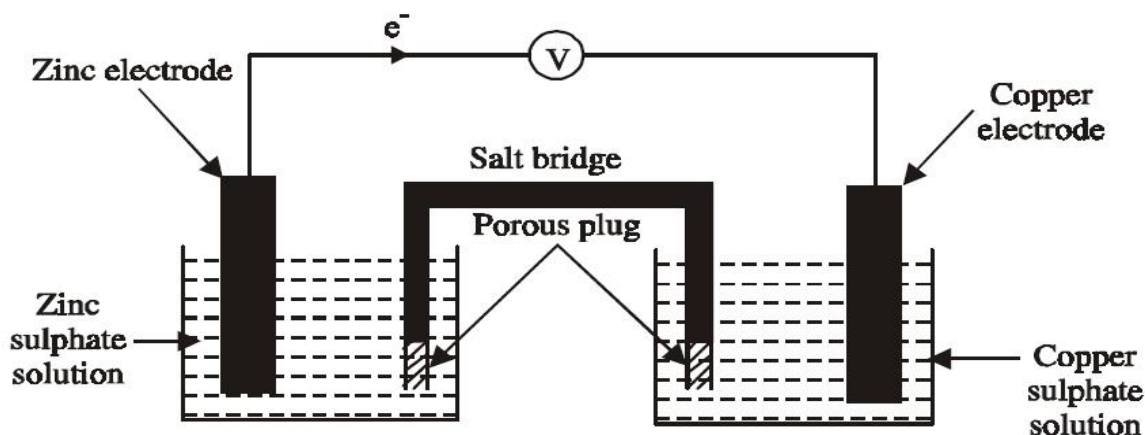
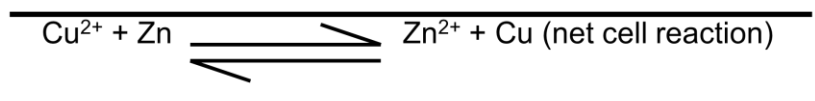
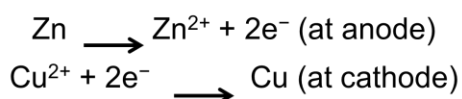


Fig. 1.7 Daniel cell

Reactions occurring in the cell

At anode: Oxidation takes place in the zinc electrode by the liberation of electrons, so this electrode is called negative electrode or anode.

At cathode: Reduction takes place in the copper electrode by the acceptance of electrons, so this electrode is called the positive electrode or cathode.



The electrons liberated by the oxidation reaction flow through the external wire and are consumed by the copper ions at the cathode.

Salt bridge

It consists of a U-tube containing saturated solution of KCl or NaNO₃ in agar-agar gel. It connects the two half cells of the galvanic cells.

Functions of salt bridge

- (i) It eliminates liquid junction potential.
- (ii) It provides the electrical continuity between the two half cells.



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